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A solution of liquid A is mixed with liquid B and exerts ideal behavior. At (103°C) the total vapor pressure of a solution is 413mmHg containing 1.4 moles of A and 2.0 moles of B. After you add another mole of A to the solution, the vapor pressure increases to 486mmHg. At (103°C) calculate the vapor pressure of pure A and B.

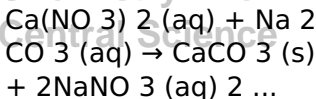
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CHEM 100: General
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9: Solutions

Expand/collapse global
location ... See
Exercise 5 and Exercise
6 above. What would
happen if the two
solutions in these
exercises were on
opposite sides of a
semipermeable
membrane, as we

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would find in our cell
walls?) Answers .



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**13.E: Solutions
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These are homework exercises to accompany Chapter 7 of the University of Kentucky's LibreText for CHE 103 - Chemistry for Allied Health. Solutions are available below the questions.

7.8: Solutions (Exercises) - Chemistry LibreTexts

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Solutions come in all phases, and the solvent and the solute do not have to be in the same phase to form a solution (such as salt and water). For example, air is a gaseous solution of about 80% nitrogen and about 20% oxygen, with some other gases present in much smaller amounts.

9.2: Solutions - Chemistry

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NCERT EXERCISES. 2.1

Define the term solution. How many types of solutions are formed? Write briefly about each type with an example. Sol. A solution is a homogeneous mixture of two or more chemically non-reacting substances.

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